

NAME \_\_\_\_\_ DATE \_\_\_\_\_ CLASS \_\_\_\_\_

### pH and pOH Calculations Worksheet

Find the pH and the pOH of the solution.

	<u>pH</u>	<u>pOH</u>
1. A 0.023 M solution of perchloric acid		
2. A $6.6 \times 10^{-6}$ M solution of nitric acid		
3. A 0.0334 M solution of potassium hydroxide		
4. A $1 \times 10^{-5}$ M solution of hydrobromic acid		
5. A $2.23 \times 10^{-2}$ M solution of calcium hydroxide		
6. A 0.000901 M solution of rubidium hydroxide		
7. A $7.23 \times 10^{-5}$ M solution of hydrochloric acid		
8. A $2.58 \times 10^{-6}$ M solution of chloric acid		
9. A 0.000469 M solution of sodium hydroxide		
10. A 0.000896 M solution of sulfuric acid		

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Find the  $[H_3O^+]$  and the  $[OH^-]$  from the following values.

	<u><math>[H_3O^+]</math></u>	<u><math>[OH^-]</math></u>
11. pH = 2.5		
12. pH = 11.3		
13. pOH = 4.6		
14. pOH = 8.7		
15. pH = 7.65		
16. pH = 6		

Determine the pH of each solution of the following bases.

17. 0.01 M sodium hydroxide \_\_\_\_\_

18. 0.001 M strontium hydroxide \_\_\_\_\_

19. 0.005 M potassium hydroxide \_\_\_\_\_

20. 0.000675 M calcium hydroxide \_\_\_\_\_